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SERIES - 1

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**Important Questions
for Chhattisgarh
Board Exam ‘2021’**

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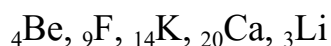
CLASS – 10th

Subject - Science

“Periodic Classification of Elements”

2 Marks Questions.

1. According to modern periodic laws properties of elements are periodic function of what?
2. How does ionization energy / valency changes in group?
3. Physical and chemical properties of elements present in a group are same ,why ?
4. What is a group / period?
5. For the classification of elements atomic number is better than atomic weight, why.
6. Choose the elements from the following which has one electron in outermost orbit and also write their electronic configuration –



7. What are metalloids ? give an example.

3 Marks Questions-

8. Write the advantages of modern periodic table.
9. In which group and which period should hydrogen be placed ? Give reason .
10. Explain Doberreiner’s law of traids?
11. Explain Newland’s law of octaves. Write its limitations.
12. Why do halogens elements posses maximum electron affinity?
13. Electronic configuration of an element is 2,8,1.
 - a. Write name of the element .
 - b. What will be the valency of the elements .
 - c. It is the element of which group ?
14. Atomic number of an element is 18 .
 - a. Write electronic configuration of the elements .
 - b. In which group will the element be placed .
 - c. What will be the name of the element ?

15. Electronic configuration of an element is 2, 8, 7.

- a. What will be the valency of the element ?
- b. It is the element of which group ?
- c. In which period will their element be placed?

4 Marks Question-

16. What is the reason behind keeping inert gases in a separate group ? Write.

17. How does electronic affinity differ from electronegativity?

18. Choose the element from following pairs which has big size and write the reason also:

- a. Mg (atomic number 12) or Cl (atomic number 17).
- b. Na (atomic number 11) or K (atomic number 19).

19. Three elements having atomic number 5, 7, and 10 respectively, state that:

- a. Which element is from group 18?
- b. Which element is from group 15?
- c. Which element is from group 13?
- d. All these elements are from which period?

5 Marks Question-

20. Compare the arrangement of elements in modern periodic table and Mendeleev periodic table.